

Chapter 10 States Of Matter Test

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[CHAPTER 10 REVIEW States of Matter SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. a When a substance in a closed system undergoes a phase change and the system reaches equilibrium, \(a\) the two opposing changes occur at equal rates. \(b\) there are no more phase changes. \(c\) one phase change predominates.](#)

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[Modern Chemistry. Chapter 10. States of Matter. Kinetic-Molecular Theory](#) [The kinetic-molecular theory of matter is based on the idea that particles of matter are always in motion and this motion has consequences that affect its physical properties. Kinetic-Molecular Theory of Gases 1- Gases consist of large numbers of tiny particles that are far apart relative to their size. 2- Collisions between gas particles and between particles and their container walls are elastic.](#)

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[Almost all substances can be classified into three states of matter – solids, liquids and gases. Each state has different properties. Heating and cooling a substance can cause it to change state.](#)

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[CHAPTER 10 States Matter - phillipsmathandphysics.com](#) [Chapter 10 – states of matter 1. Chapter 10: States of Matter• Kinetic Molecular Theory: the idea that particles of matter are always in motion; this applies to all states of matter• Used to explain properties of matter in terms of energy 2.](#)

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[CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided](#) [Modern chemistry chapter 10 review answers states of matter section 1. 1. Match description on the right to the correct crystal type on the left](#) [Modern chemistry chapter 10 review answers states of matter section 1. b ionic crystal \(a\) has mobile electrons in the crystal c covalent ...](#)

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[CHEMISTRY NOTES – Chapter 10 States of Matter Goals : To gain an understanding of : 1.. The kinetic theory of matter. 2. Avogadro's hypothesis. 3. The nature of liquids, solids and gases. 4. Phase changes. NOTES: The kinetic theory of matter states that particles which make up all types of matter are in constant motion. CHEMISTRY NOTES Chapter 10 States of Matter](#)