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CHAPTER 1 The Basics of Reaction

Kinetics for Chemical Reaction

Engineering. The definition of the

activity of species i is: fugacity at the
standard state, that is, 1 atm for gases

and thus. 3. $K = \frac{f_i}{f_i^\ominus} = \frac{f_i}{1 \text{ atm}}$

Use of the

Lewis and Randall rule gives:

$f_i = \gamma_i P_i$ at 1 atm.

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POWER LAW BASICS | 6 2 POWER LAW BASICS 2.1 ARRHENIUS EQUATION

The main assumption behind the Arrhenius expression is that $k = A e^{-E_a/RT}$ (Equation 2.1). This is an approximation, but it works quite well. The rate coefficient is the term that is a function of temperature but may also

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Chemical Equilibrium $dG = -SdT + VdP + \sum_{j=1}^n \mu_j dn_j$ μ_j : chemical potential for species j . $G_{i,T,P} = \sum_{j=1}^n X_j \mu_j = 0, i = 1, \dots, n$ $\mu_j = G_j + RT \ln a_j, a_j = f_j / f_j^K$ $i = \sum_{j=1}^n Y_j = 1$ $a_{ij} = G_i = -RT \ln K_i$ Standard state: pure species j at 1 atm and system temperature.

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Description. Chemical kinetics is the study of the rates of chemical reactions; if changes in conditions impact the speed of a reaction, we can better understand what caused the reaction. In this course, you ' ll learn to understand complicated data sets and analysis techniques for measuring and understanding these rate changes.

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